Date Planned : __ / __ / __	Daily Tutorial Sheet-1	Expected Duration : 45 Min
Actual Date of Attempt : __ / __ / __	JEE Advanced Archive	Exact Duration :_____

1. An isomer of ethanol : **(1980)**

(A) methanol **(B)** diethyl ether **(C)** acetone **(D)** dimethyl ether

2. Which of the following will have least hindered rotation about carbon-carbon bond ? **(1981)**

(A) Ethane **(B)** Ethylene **(C)** Acetylene **(D)** Hexachloroethane

3. The compound which is not isomeric with diethyl ether is : **(1981)**

(A) *n*-propyl methyl ether **(B)** butan-1-ol

(C) 2-methylpropane-2-ol **(D)** butanone

4. ..... ring is most strained. (cyclopropane, cyclobutane, cyclopentane) **(1981)**

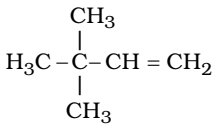
5. The compound having both sp and sp^2-hybridised carbon atoms is ..... (propane, propene, propadiene).

6. Among the given cations, ... is most stable. **(1981)**
(sec-butyl carbonium ion, tert-butyl carbonium ion, n-butyl carbonium ion)

7. Which of the following compounds will exhibit cis-trans (geometrical) isomerism ? **(1983)**

(A) 2-butene **(B)** 2-butyne **(C)** 2-butanol **(D)** butanol

8. The IUPAC name of the compound having the formula is : **(1984)**



(A) 3,3, 3-trimethyl-1-propene **(B)** 1, 1, 1-trimethyl-2-propene

(C) 3, 3-dimethyl-1-butene **(D)** 2, 2-dimethyl-3-butene

9. Write structural formulae for the isomeric alcohols having the molecular formula $C_4H_{10}O$? **(1984)**

10. The terminal carbon atom in butane is..... hybridized. **(1985)**

11. m-chlorobromobenzene is an isomer of m-bromochloro benzene. (T/F) **(1985)**

12. Write the structure of all the possible isomers of dichloroethene. Which of them will have zero dipole moment ? **(1985)**

13. A .... diol has two hydroxyl groups on ..... carbon atoms. **(1985)**

14. The IUPAC name of the compound $CH_2 = CH - CH(CH_3)_2$ is : **(1987)**

(A) 1-1-dimethyl-2-butene **(B)** 3-methyl-1-butene

(C) 2-vinyl propane **(D)** None of these

15. The bond between carbon atom (1) and carbon atom (2) in compound $N \equiv \underset{1}{C} - \underset{2}{C}H = CH_2$ Involves the hybridization as : **(1987)**

(A) sp^2 and sp^2 **(B)** sp^3 and sp **(C)** sp and sp^2 **(D)** sp and sp